

Course Prefix/Number/Title: Chem. 115, Introductory Chemistry

Number of Credits: 4

Course Description: The topics covered will include measurements, ionic and covalent compounds, chemical calculations, states of matter, energy, solutions, reactions and chemical bonding. The course is designed for non-science majors and students in the nursing program.

Pre-/Co-requisites: ASC 93

Course Objectives:

1. Students will gain an understanding of the nature of atoms, molecules, elements, chemical bonds and compounds.
2. Students will gain a basic understanding of the changes that take place in chemical reactions. Ability to perform simple stoichiometry calculations.
3. Students will gain an understanding of the phases of matter.
4. Students will gain an understanding and use of scientific methods
5. Students will gain an elementary understanding of the nature of acids and bases.

Instructor: Angie Bartholomay

Office: NSC 111

Office Hours: MW 9:00-10:00am, MTF 1:00-2:00pm

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Lecture/Lab Schedule: Lecture MWF 11:00-11:44am lab T 8:00-9:50am

Textbook(s): Introductory Chemistry, by Zumdahl, 6th edition.

Course Requirements: Grading: Grades will be based on total points using the following percentage system: 100-90, A; 89-80, B; 79-70, C; 69-60, D; <60, F. Exams, research paper, and homework quizzes, and lab reports will be used to determine the final grade. **IMPORTANT!** Any grievances concerning graded material must be addressed within one week from the time the material is returned to the student.

Exams (5)	500pts
Lab Reports (25 pts. Each)	300pts
Final Lab	100pts
Quizzes (10pts. Each)	<u>100pts</u>
	1000pts

Exams: There will be five exams during the course of the semester. Exams may contain short answer/essay, multiple choice, completion and problems. There will be no makeup exams unless prior arrangements have been made. If you need to be gone for a school related activity or family event, you will be expected make arrangement prior to the event and take the exam before you leave.

Homework: Homework will be assigned throughout the semester and will be discussed in class, these assignments can be used on quizzes. Homework is designed to prepare you for exams and quizzes. These quizzes will be unannounced throughout the semester there will not be make-up quizzes. You are expected to read the assigned pages prior to class. Lecture may not cover everything assigned in the reading, but everything assigned is exam material. If you do not understand something in the readings, please ask questions.

Laboratory: The laboratory portion of the course provides an opportunity to integrate lecture concepts with observable activities. Attendance at lab is mandatory! **Chemical splash safety goggles** and metric ruler are required and may be purchased at the bookstore. Failure to wear goggles will result in a reduction in lab report grades and continued omission will result in removal from lab activities and a loss of all remaining lab points available. To obtain credit, you must be actively involved in the laboratory activities and is turned in at the end of the period.

Early Warning Attendance Policy will be followed

Tentative Course Outline:

<u>Lecture</u>	<u>Chapter and Reading Assignment</u>	<u>Lab Topic</u>
Week 1	Ch. 1-2, Pages 1-18 scientific method	No Lab
	Ch. 2, Pages 18-33 measurements & calculations	
Week 2	Ch. 2-3, Pages 33-66 matter	measurement, accuracy, density
	Ch. 3, wrap-up and review	
	<u>Chapter #1-3 Exam</u>	
Week 3	Ch. 4, p. 72-88 Elements, atoms, ions	percent composition
	Ch. 4, pages 89-104	
Week 4	Ch. 5, Pages 112-126 nomenclature	
	Ch. 5&6, Pages 126-149	Chemical reaction physical & chemical change
Week 5	Ch. 6&7, Pages 149-175	
	Ch. 7, Pages 175-191 aqueous reactions	empirical formulas
	Ch. 4-7 wrap up and review	
	<u>Ch. 4-7 Exam</u>	
Week 6	Ch. 8, Pages 203-218	Chemical composition
	Ch. 8, Pages 218-229	chemical reactions
Week 7	Ch. 9, Pages 239-251	Chemical quantities relating moles to coefficients in equations
	Ch. 9, Pages 251-259	
Week 8	<u>Ch. 8-9 Exam</u>	mole & mass relationships
	Ch. 10, Pages 271-286	Energy
Week 9	Ch. 10, Pages 287-297	calorimetry
	Ch. 11, Pages 303-316	
Week 10	Ch. 11, Pages 317-332	Fame tests
	Ch. 12, Pages 341-356	chemical bonding
Week 11	Ch. 12, Pages 356-373	molecular geometry and valence electrons
	<u>Ch. 10-12 Exam</u>	
Week 12	Ch. 14, pages 427-432	phases of matter
	Ch. 14, Pages 432-444	
Week 13	Ch. 15, Pages 451-462	solutions
	Ch. 15, Pages 462-473	solubility of a salt
Week 14	<u>Ch. 13-15 Exam</u>	
	Ch. 16, Pages 487-507	acids & bases
Week 15	Ch. 17, Pages 515-526	properties of acids & Bases
	Ch. 17, Pages 526-541	equilibrium
Week 16	Ch. 18, Pages 553-575	redox reactions
	<u>Ch. 16-18 Exam</u>	

General Education Competency/Learning Outcome(s) OR CTE Competency/Department Learning Outcome(s): General Education Competency #1: Identifies the interrelationships between humans and their environment.

Learning Outcomes #1- Applies scientific methods of inquiry

Learning Outcomes #3- Applies scientific information in everyday life

Relationship to Campus Focus: This course addresses the campus theme by incorporating the role that chemistry plays in our everyday life and the impact it has on our natural world. In addition students will use technology to conduct labs as well as study how technology can be used in chemistry. The course will address the role of chemistry in their everyday life as well as in their future.

Classroom Policies: make-up for missed exams will not be allowed unless prior arrangements have been made. If you must be absent for a school related or family event, you are expected to make prior arrangements and take the exam prior to the event. If you are given permission to take a late exam you will have 48 hours to make it up.

No electronic devices will be allowed. Cell phone must be turned off at all times in class! You will be asked once to put the phone away, if asked again you will be asked to leave.

Headphones will not be allowed during exams!

Be respectful of other students, instructors, and guests!

Student Email Policy:

Dakota College at Bottineau is increasingly dependent upon email as an official form of communication. A student's campus-assigned email address will be the only one recognized by the Campus for official mailings. The liability for missing or not acting upon important information conveyed via campus email rests with the student.

Academic Integrity:

According to the DCB Student Handbook, students are responsible for submitting their own work. Students who cooperate on oral or written examinations or work without authorization share the responsibility for violation of academic principles, and the students are subject to disciplinary action even when one of the students is not enrolled in the course where the violation occurred. The Code detailed in the Academic Honesty/Dishonesty section of the Student Handbook will serve as the guideline for cases where cheating, plagiarism or other academic improprieties have occurred.

Disabilities or Special Needs:

Students with disabilities or special needs (academic or otherwise) are encouraged to contact the instructor and Disability Support Services.

Title IX:

Dakota College at Bottineau (DCB) faculty are committed to helping create a safe learning environment for all students and for the College as a whole. Please be aware that all DCB employees (other than those designated as confidential resources such as advocates, counselors, clergy and healthcare providers) are required to report information about such discrimination and harassment to the College Title IX Coordinator. This means that if a student tells a faculty member about a situation of sexual harassment or sexual violence, or other related misconduct, the faculty member must share that information with the College's Title IX Coordinator. Students wishing to speak to a confidential employee who does not have this reporting responsibility can find a list of resources on the DCB Title IX webpage.